

TABLE 16.2 Solubility Products of Some Slightly Soluble Ionic Compounds at 25°C

Compound	K_{sp}	Compound	K_{sp}
Aluminum hydroxide [Al(OH) ₃]	1.8×10^{-33}	Lead(II) chromate (PbCrO ₄)	2.0×10^{-14}
Barium carbonate (BaCO ₃)	8.1×10^{-9}	Lead(II) fluoride (PbF ₂)	4.1×10^{-8}
Barium fluoride (BaF ₂)	1.7×10^{-6}	Lead(II) iodide (PbI ₂)	1.4×10^{-8}
Barium sulfate (BaSO ₄)	1.1×10^{-10}	Lead(II) sulfide (PbS)	3.4×10^{-28}
Bismuth sulfide (Bi ₂ S ₃)	1.6×10^{-72}	Magnesium carbonate (MgCO ₃)	4.0×10^{-5}
Cadmium sulfide (CdS)	8.0×10^{-28}	Magnesium hydroxide [Mg(OH) ₂]	1.2×10^{-11}
Calcium carbonate (CaCO ₃)	8.7×10^{-9}	Manganese(II) sulfide (MnS)	3.0×10^{-14}
Calcium fluoride (CaF ₂)	4.0×10^{-11}	Mercury(I) chloride (Hg ₂ Cl ₂)	3.5×10^{-18}
Calcium hydroxide [Ca(OH) ₂]	8.0×10^{-6}	Mercury(II) sulfide (HgS)	4.0×10^{-54}
Calcium phosphate [Ca ₃ (PO ₄) ₂]	1.2×10^{-26}	Nickel(II) sulfide (NiS)	1.4×10^{-24}
Chromium(III) hydroxide [Cr(OH) ₃]	3.0×10^{-29}	Silver bromide (AgBr)	7.7×10^{-13}
Cobalt(II) sulfide (CoS)	4.0×10^{-21}	Silver carbonate (Ag ₂ CO ₃)	8.1×10^{-12}
Copper(I) bromide (CuBr)	4.2×10^{-8}	Silver chloride (AgCl)	1.6×10^{-10}
Copper(I) iodide (CuI)	5.1×10^{-12}	Silver iodide (AgI)	8.3×10^{-17}
Copper(II) hydroxide [Cu(OH) ₂]	2.2×10^{-20}	Silver sulfate (Ag ₂ SO ₄)	1.4×10^{-5}
Copper(II) sulfide (CuS)	6.0×10^{-37}	Silver sulfide (Ag ₂ S)	6.0×10^{-51}
Iron(II) hydroxide [Fe(OH) ₂]	1.6×10^{-14}	Strontium carbonate (SrCO ₃)	1.6×10^{-9}
Iron(III) hydroxide [Fe(OH) ₃]	1.1×10^{-36}	Strontium sulfate (SrSO ₄)	3.8×10^{-7}
Iron(II) sulfide (FeS)	6.0×10^{-19}	Tin(II) sulfide (SnS)	1.0×10^{-26}
Lead(II) carbonate (PbCO ₃)	3.3×10^{-14}	Zinc hydroxide [Zn(OH) ₂]	1.8×10^{-14}
Lead(II) chloride (PbCl ₂)	2.4×10^{-4}	Zinc sulfide (ZnS)	3.0×10^{-23}